

Monoatomic ions

Chemical formula	Name
H ⁺	hydrogen ion
Li ⁺	lithium ion
Na ⁺	sodium ion
K ⁺	potassium ion
Mg ²⁺	magnesium ion
Ca ²⁺	calcium ion
Sr ²⁺	strontium ion
Ba ²⁺	barium ion
Al ³⁺	aluminum ion
H ⁻	hydride ion
F ⁻	fluoride ion
Cl ⁻	chloride ion
Br ⁻	bromide ion
I ⁻	iodide ion
O ²⁻	oxide ion
S ²⁻	sulfide ion
N ³⁻	nitride ion

Transition metal cations

Atom	Charge(s)	Name
silver (Ag)	+1	silver
cobalt (Co)	+2	cobalt (II), cobaltous
	+3	cobalt (III), cobaltic
copper (Cu)	+1	copper (I), cuprous
	+2	copper (II), cupric
cadmium (Cd)	+2	cadmium
chromium (Cr)	+2	chromium (II), chromous
	+3	chromium (III), chromic
iron (Fe)	+2	iron (II), ferrous
	+3	iron (III), ferric
mercury (Hg)	+1	mercury (I), mercurous
	+2	mercury (II), mercuric
manganese (Mn)	+2	manganese (II), manganous
	+3	manganese (III), manganic
nickel (Ni)	+2	nickel (II), nickelous
	+3	nickel (III), nickelic
lead (Pb)	+2	lead (II), plumbous
	+4	lead (VI), plumbic
tin (Sn)	+2	tin (II), stannous
	+4	tin (VI), stannic
zinc (Zn)	+2	zinc

Note that the mercury (I)/mercurous cation is Hg_2^{2+} , where each atom is given a +1 charge.

Oxyanions

Chemical formula	Name
ClO^-	hypochlorite
ClO_2^-	chlorite
ClO_3^-	chlorate
ClO_4^-	perchlorate
BrO^-	hypobromite
BrO_2^-	bromite
BrO_3^-	bromate
BrO_4^-	perbromate
IO^-	hypoiodite
IO_2^-	iodite
IO_3^-	iodate
IO_4^-	periodate
SO_3^{2-}	sulfite
SO_4^{2-}	sulfate
HSO_4^-	hydrogen sulfate, bisulfate
$\text{S}_2\text{O}_3^{2-}$	thiosulfate
PO_3^{3-}	phosphite
PO_4^{3-}	phosphate
NO_2^-	nitrite
NO_3^-	nitrate
AsO_3^{3-}	arsenite
AsO_4^{3-}	arsenate
CO_3^{2-}	carbonate
HCO_3^-	hydrogen carbonate, bicarbonate
$\text{C}_2\text{O}_4^{2-}$	oxalate
CrO_4^{2-}	chromate
$\text{Cr}_2\text{O}_7^{2-}$	dichromate
MnO_4^-	permanganate
$\text{C}_2\text{H}_3\text{O}_3^-$	acetate
CNO^-	cyanate
CNS^-	thiocyanate

Other common ions

Chemical formula	Name
NH_4^+	ammonium
H_3O^+	hydronium
N_3^-	azide
CN^-	cyanide
HS^-	hydrogen sulfide
OH^-	hydroxide
O_2^{2-}	peroxide